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Solution: (a)  $(1000 \times 10^3) (6.022 \times 10^{23}$   
atoms/mol)  $(55.847 \text{ g/mol})$  (b)  $(1 \text{ mol})$   
 $(10.81 \text{ g/mol}) = 4.6 \text{ cm}^3$   $2.36 \text{ g/cm}^3$   
 $= 1.078 \times 10^{28}$  atoms

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